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Synergetic effect of CuS@ZnS nanostructures on photocatalytic degradation of organic pollutant under visible light irradiation†

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Ultrafast visible light active CuS/ZnS nanostructured photocatalysts were synthesized by a hydrothermal method. The effect of the CuS concentration on the morphological, structural and optical properties of ZnS nanostructures were investigated. X-ray diffraction analysis indicated the formation of CuS/ZnS phases with good crystallinity. The presence of ZnS on CuS was confirmed by X-ray photoelectron spectroscopy, elemental mapping, scanning electron microscopy and high resolution transmission electron microscopy analyses. CuS/ZnS nanocomposites efficiently decomposed methylene blue (MB) upon irradiation with visible light. The degradation time was 3 times faster than that of bare ZnS and CuS. The scavenger analysis results revealed that photogenerated superoxide ($O_2^{\cdot-}$) radicals were the main reactive species for the degradation of MB. The maximum degradation efficiency of 95.51% was observed within 5 min for samples of 0.050 M CuS/ZnS composites. The obtained results suggested that the enhanced photocatalytic activity was due to the formation of CuS/ZnS heterojunctions which led to the efficient separation of the photoinduced carriers.

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1. Introduction

Photocatalysts have been explored for their utility in environmental applications such as pollutant dissociation, electric energy conversion devices, and water-splitting applications which are required for the quality of our life.^{1–6} Photocatalysis provides a potential strategy to decompose contaminants by light-driven chemical processes over the surface of the photocatalyst. Honda and Fujishima's pioneering work of photocatalytic water splitting on TiO_2 electrodes has attracted widespread attention in clean, environmentally friendly and economical processes.^{7,8} Various semiconductor photocatalysts, including metal oxides, metal hydroxides, oxynitrides, sulfides and metal-free semiconductors, have been investigated for water purification.^{9–12} Among the photocatalysts, zinc sulfide (ZnS) has been studied extensively by several researchers.¹³ ZnS is an important II–VI group

semiconductor, which can rapidly produce electron–hole pairs by photoexcitation and thus create high negative potentials of the excited electron. The theoretical efficiency of the photo-carrier generation of ZnS is much higher than that of TiO_2 because of a direct band gap.¹⁴ However, the efficiency of ZnS is impeded by a wide band gap energy ($E_g = 3.66$ eV) that strictly precludes the utilization of visible light and results in a reduced viability of ZnS as a visible light photocatalyst. These photocatalysts are mostly active only under UV light, which only occupies 4% of the total sunlight, and thus restricts its practical application.¹⁵

A composite with two semiconductors having a p–n junction can induce the transition of the optical responsibility from UV to the visible region and separate the electron–hole pairs.^{16,17} Thus, an improved charge separation, increased charge carrier lifetime and efficient interfacial charge transfer from the catalyst to the adsorbed substance are achieved in the presence of composite semiconductors as the photocatalyst, resulting in an enhanced photocatalytic activity. It is known that the photocatalytic activity of a photocatalyst mainly depends on an amount of the photo-excited active species and the separation efficiency of the photo-excited electron–hole pairs.¹⁸ The formation of the p–n junction in a photocatalytic system provides an internal electrostatic field with a direction from the n-type semiconductor to the p-type semiconductor. The photoexcited electron–hole pairs would be thus effectively separated by moving the electron and hole in opposite directions and their recombination can be minimized.^{19–21} Therefore, it is desirable to design an efficient visible

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light driven photocatalyst in which one component is a p-type semiconductor and the other is an n-type semiconductor.

Composite nanostructures have been studied for these photocatalytic applications. For example, Lee *et al.* fabricated CuS (p-type) and ZnO (n-type) nanoarrays on a stainless steel mesh and applied them as a visible light photocatalyst for the decomposition of acid orange 7.²² Zhang *et al.* showed that p-type NiO and n-type ZnO heterojunction nanofibres exhibited a superior photocatalytic activity over pure NiO and ZnO nanofibres for the degradation of Rhodamine B. The separation of photogenerated electron-hole pairs was highly enhanced.²³ Soltani *et al.* synthesized ZnS and CdS nanoparticles which extended the light absorption in the visible region and showed an increased stability against photo-corrosion.²⁴ Babu *et al.* synthesized Cu/TiO₂ core-shell nanowires with 3% Cu/TiO₂ and indicated the highest degradation rate of methyl orange.²⁵ Sunita *et al.* synthesized TiO₂/CuS core-shell nanostructures. Methylene blue was degraded within 60 min of irradiation.²⁶ Basu *et al.* prepared ZnO/CuS heterostructure by decorating CuS nanostructures on the surface of ZnO nanotubes by a wet chemical method. They showed that the ZnO/CuS heterostructure efficiently decomposed methylene blue (MB) upon irradiation of visible light at room temperature.²⁷ The composites facilitated charge transfer across heterojunctions, effectively improving the separation of photogenerated electron-hole pairs.²⁸ The above research clearly demonstrated that the coupling of semiconductors significantly enhances their photocatalytic activities. Copper sulfide (CuS) is a p-type semiconductor with a narrow band gap ($E_g = 2.2$ eV) which could be a highly efficient and stable visible-light photocatalyst for pollutant treatment.^{29–31} However, the rapid recombination of the photogenerated electron-hole pairs of CuS restricts the improvement of its degradation efficiency as a photocatalyst. The formation of a heterojunction between ZnS and CuS in the nanocomposite photocatalyst is expected to promote the transfer of photogenerated electrons and holes between ZnS and CuS, which suppresses their recombination and enhances the photocatalytic activity.^{32–35} Therefore, CuS/ZnS nanocomposites can be a promising material as a photocatalyst because of their physical and chemical properties.

In the present work, a novel approach is reported for the formation of well-connected ZnS/CuS nanostructures by

a hydrothermal method. The effect of CuS concentration on the phase and morphology of the composites is investigated. The functional properties of the ZnS/CuS nanocomposites are studied by X-ray diffraction (XRD), field-emission scanning electron microscopy (FESEM), transmission electron microscopy (TEM) and X-ray photoelectron spectroscopy (XPS). The photocatalytic activity of the synthesized materials is obtained by quantifying the rate of MB degradation in an aqueous suspension under visible light irradiation. The kinetics and photocatalytic mechanism of ZnS/CuS nanocomposites are proposed.

2. Experimental procedure

All the chemicals were purchased from Wako chemicals (Japan) and used without further purification. In a typical synthesis,

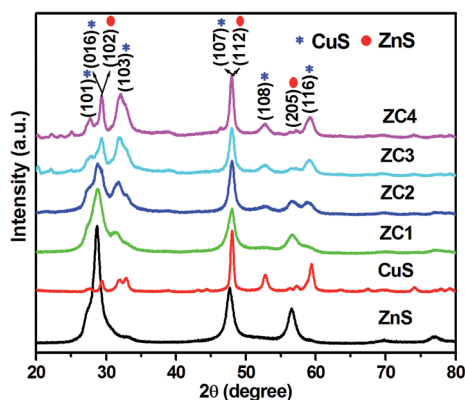


Fig. 1 XRD patterns of ZnS, CuS and CuS/ZnS nanocomposites.

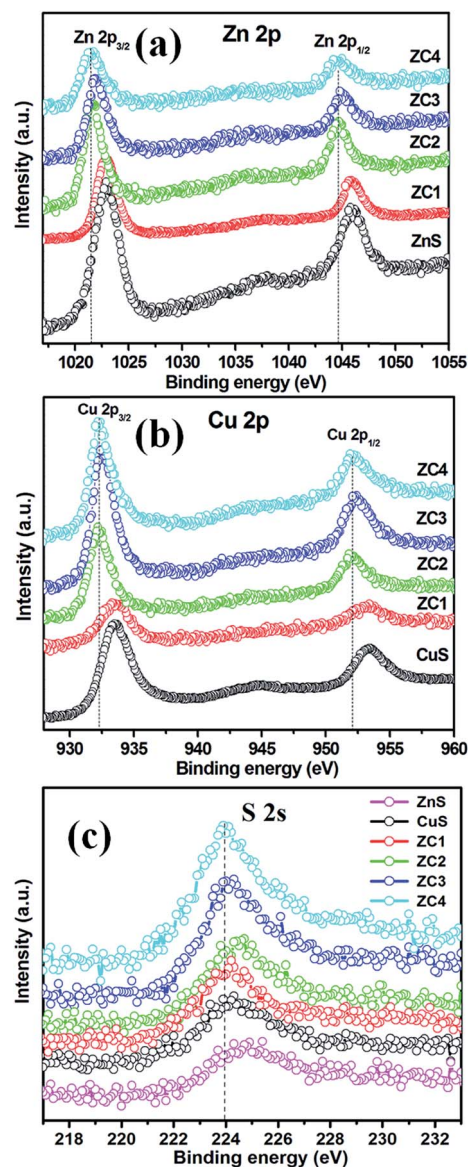


Fig. 2 XPS spectra of (a) Zn 2p, (b) Cu 2p and (c) S 2s states.



0.1 mol L⁻¹ of copper acetate was dissolved in 100 mL of deionized water and allowed to stir until complete dissolution of the compound. 1 mL of triethylamine (TEA) was added to the above solution as a stabilizing agent. 0.1 mol L⁻¹ of thioacetamide was added and allowed to stir for 12 h. Finally, the mixture was transferred to a Teflon-lined autoclave and maintained at 150 °C for 12 h in a hot air oven. The obtained precipitate was separated by centrifugation, and washed with deionized water and ethanol several times. The samples were termed CuS. For the synthesis of ZnS nanostructures, the same procedure was followed with a small modification. 0.1 mol L⁻¹ of zinc acetate was added instead of copper acetate and the sample was termed ZnS.

2.1. Synthesis of CuS/ZnS nanocomposites

0.1 mol L⁻¹ of zinc acetate was dissolved in 100 mL of deionized water and allowed to stir until complete dissolution of the compound. Copper acetate of different mol percentages (0.025, 0.050, 0.075 and 0.1 M) was added to the zinc acetate solution under continuous stirring. 1 mL of TEA and 0.1 mol L⁻¹ of thioacetamide were added and allowed to stir for 12 h. Finally, the mixture was transferred to a Teflon-lined autoclave and maintained at 150 °C for 12 h in a hot air oven. The obtained product was separated by centrifugation, and washed with deionized water and ethanol several times. The obtained product was dried at 80 °C for 1 h. The samples were termed as ZC1, ZC2, ZC3 and ZC4 for the copper acetate concentrations of 0.025, 0.050, 0.075 and 0.1 M, respectively.

2.2. Characterization techniques

The obtained products were characterized by XRD using a Rigaku X-ray diffractometer (RINT 2200, Japan) with CuK α radiation and 0.02° s⁻¹ step interval. The morphology and particle size of the products were assessed by FESEM (JEOL JSM 7001F microscope) at an accelerating voltage of 15 kV and TEM (JEOL JEM 2100F microscope) at an accelerating voltage of 200 kV. UV-visible absorption spectra were performed by a Shimadzu 3101 PC spectrophotometer (Japan). XPS spectra were measured by a Shimadzu ESCA 3400.

2.3. Photocatalytic studies

Evaluation of the photocatalytic performance of CuS/ZnS composites was explored by examining the photo-assisted degradation of MB under visible light illumination [xenon light source (MAX-303, Asahi Spectra)]. The photocatalytic ability tests were accomplished in a 50 mL glass beaker. In the photocatalytic reaction, the dye concentration was fixed at 10 ppm, and the photocatalyst of a known dosage (75 mg L⁻¹) was added to the above dye solution. The suspensions were stirred to achieve an absorption-desorption equilibrium state for the solution which was required to be kept in the dark before light irradiation.^{36,37} The reaction mixture was irradiated with stirring under the xenon lamp positioned at 21 cm above the reaction mixture. A blank experiment was executed either in the absence of sample (photocatalyst) or in the dark to determine that degradation of MB is truly operated by a photocatalytic mechanism. At regular time intervals, 3 mL of the suspension was collected, centrifuged, and analysed by a UV-vis spectrometer. The MB degradation was estimated from the decrease in

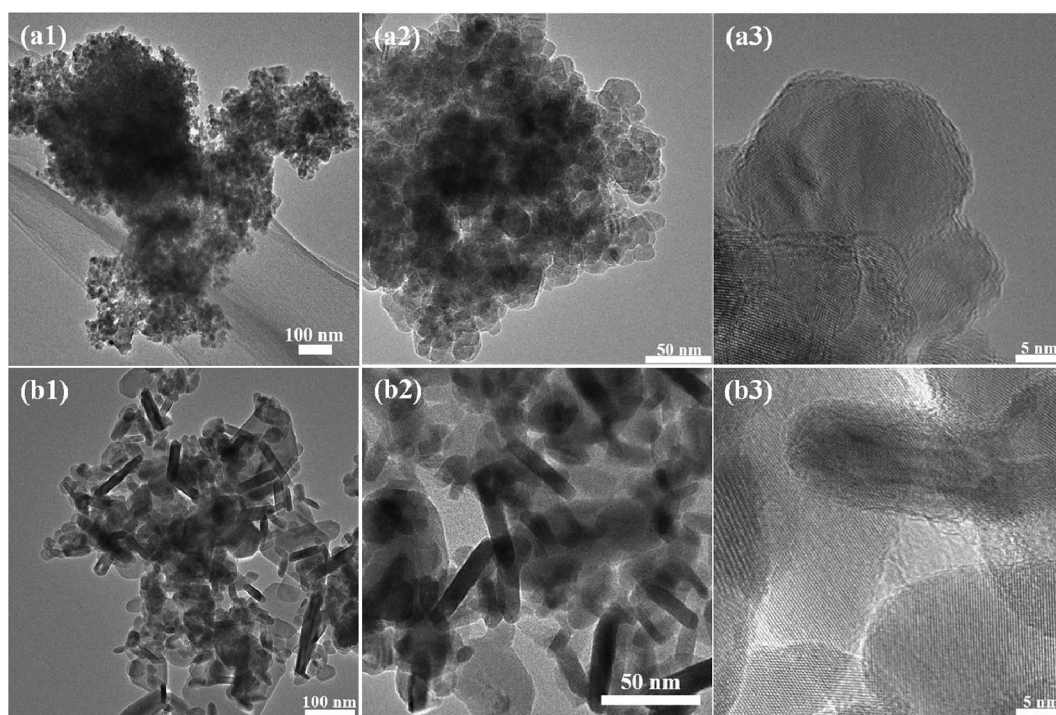


Fig. 3 TEM and HRTEM images of ZnS (a1–a3) and CuS (b1–b3).



the intensity of the associated characteristic band absorption at 664 nm. The photodegradation percentage of MB was calculated using the following equation:³⁸

$$D(\%) = \frac{C_0 - C_t}{C_0} \times 100, \quad (1)$$

where C_0 refers to the absorption of MB after the absorption equilibrium is achieved prior to the visible light illumination

and C_t is the concentration of MB at time interval t under visible light illumination.

To achieve a full understanding of the photocatalytic mechanism of MB over the CuS/ZnS composites, trapping experiments utilizing active species trappers were conducted. Different scavengers were used to trap photogenerated electrons and holes in the photocatalysis. Potassium persulfate ($K_2S_2O_8$) and benzoic acid (BA) were used to trap

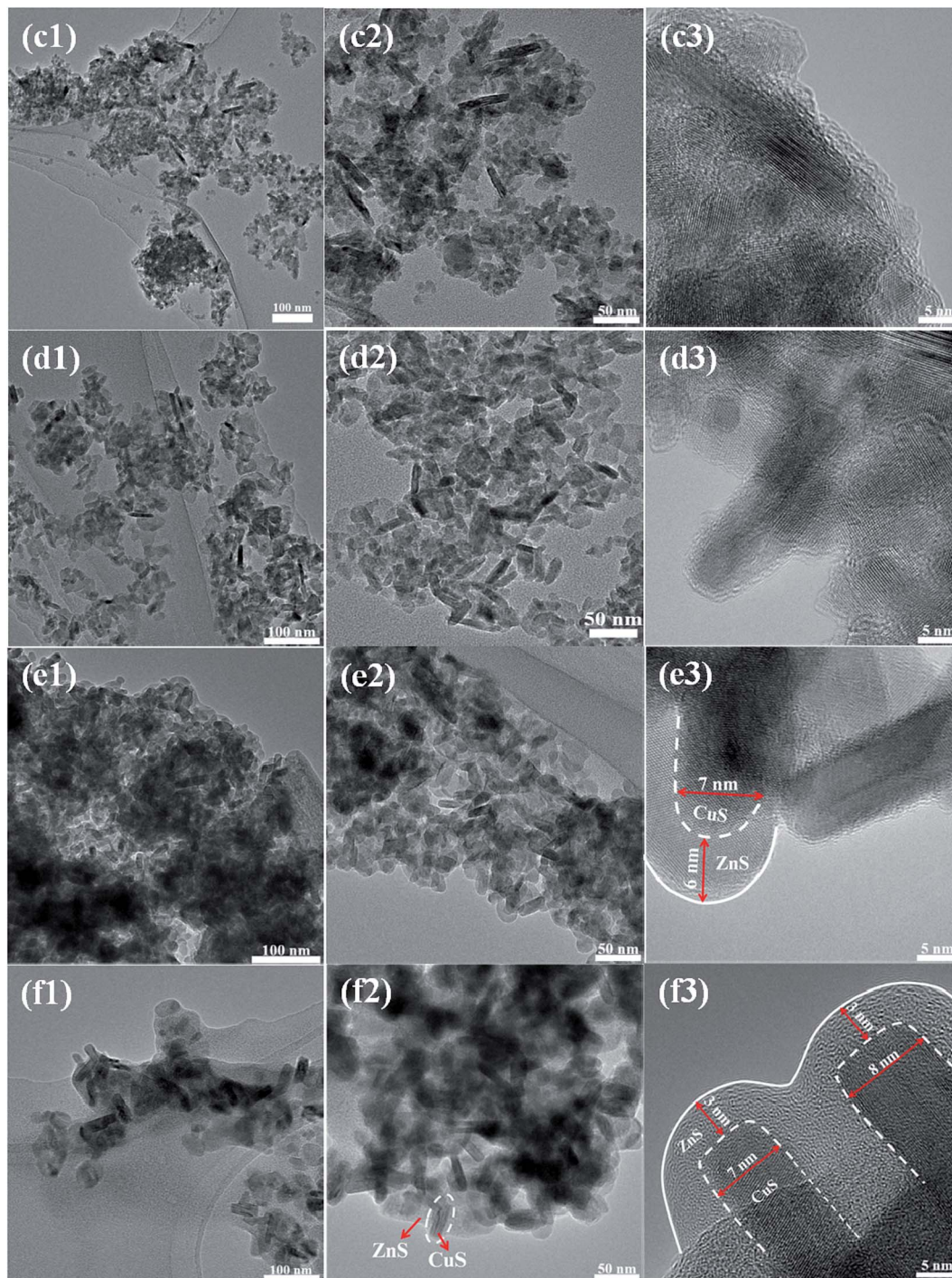


Fig. 4 TEM and HRTEM images of CuS/ZnS composites ZC1 (c1–c3), ZC2 (d1–d3), ZC3 (e1–e3) and ZC4 (f1–f3).



electrons and holes, keeping all the other conditions unaltered.

3. Results and discussion

XRD patterns of pure ZnS, CuS and CuS/ZnS composites are shown in Fig. 1. The XRD pattern of pure ZnS exhibited several diffraction peaks at 28.70° , 47.61° and 56.55° , which corresponded to the (102), (112) and (205) planes of the cubic phase of ZnS (JCPDS card no. 89-2194). CuS sample possessed diffraction peaks at 27.51° , 29.37° , 31.74° , 48.11° , 52.67° and 59.42° which corresponded to (101), (016), (103), (107), (108) and (116) planes of the hexagonal phase of CuS (JCPDS card no. 24-0060). The samples of ZC1–ZC4 showed the XRD pattern of CuS/ZnS composites. The sample ZC1 showed mixed planes that belonged to the (103) and (116) planes of CuS, whereas the samples with a higher CuS concentration, *i.e.* ZC2 to ZC4, possessed all the peaks of the cubic ZnS and hexagonal CuS patterns. This depicted the formation of ZnS/CuS composites. The intensity of the ZnS peaks decreased with an increase in the concentration of CuS.^{39,40} All the diffraction peaks indicated that the materials crystallized well as composites. The structures of ZnS and CuS in all the samples were identified as cubic (JCPDS card no. 89-2194) and hexagonal (JCPDS card no. 89-2194).⁴¹

The elemental composition and chemical states of ZnS, CuS and CuS/ZnS composite samples were analysed using XPS. The high-resolution XPS spectra of Zn 2p, Cu 2p and S 2s are shown in Fig. 2(a), (b) and (c), respectively. The core-level spectra of the Zn 2p state of pure ZnS displayed two peaks centred at 1022.97 eV and 1045.92 eV which corresponded to Zn 2p_{3/2} and Zn 2p_{1/2} suggesting the existence of Zn²⁺.^{42,43} The energy difference between the two binding energies was 22.95 eV,

which was in good agreement with the standard value.^{44,45} The peak positions of samples ZC1 were shifted from 1022.97 to 1022.87 eV and from 1045.92 to 1045.82 eV when compared to pure ZnS. When the CuS concentration increased, the peak positions were shifted to 1021.69 eV and 1044.75 eV for sample ZC2, 1021.69 eV and 1044.64 eV for sample ZC3 and 1021.48 eV and 1044.53 eV for sample ZC4. A slight shift in the binding energies of the CuS/ZnS composite samples was observed due to the electronic exchange caused by the difference the electro-negativity of zinc (1.6) and copper (1.9) elements.⁴⁶

As shown in Fig. 2(b), two typical peaks at 933.52 eV and 953.44 eV corresponded to Cu 2p_{3/2} and Cu 2p_{1/2}, respectively. These peaks have the same binding energies as the Cu 2p orbital of the Cu²⁺ state.^{47,48} The peak positions of sample ZC1 were shifted from 933.52 eV to 933.34 eV and from 953.44 eV to 953.26 eV when compared to pure CuS. Similar peak shifts were observed in the binding energies of CuS/ZnS composites from 933.34 to 932.26 eV and from 953.26 to 952.10 eV, respectively. The binding energy of the S 2s spectra centred at 224.69 eV is the characteristic value for metal sulfides (Fig. 2(c)).⁴⁹ A similar peak shift from 224.69 eV to 223.98 eV was observed in the composite samples. By the above analyses, it can be concluded that there is a strong interaction between ZnS and CuS in the composites.

The morphology of the sample was observed by FESEM and TEM analysis, as shown in Fig. 3 and 4. Fig. 3(a1)–(a3) shows the FESEM and TEM images of pure ZnS. The pure ZnS possessed spherical nanoparticles. All the particles were in the range of 10–20 nm in diameter with crystalline fringes, as observed by HRTEM in Fig. 3(a3). Whereas the pure CuS (Fig. 3(b1)–(b3)) exhibited a morphology of nanorods. The size of the nanorods was in the range of about 15–20 nm in width and 100–300 nm in

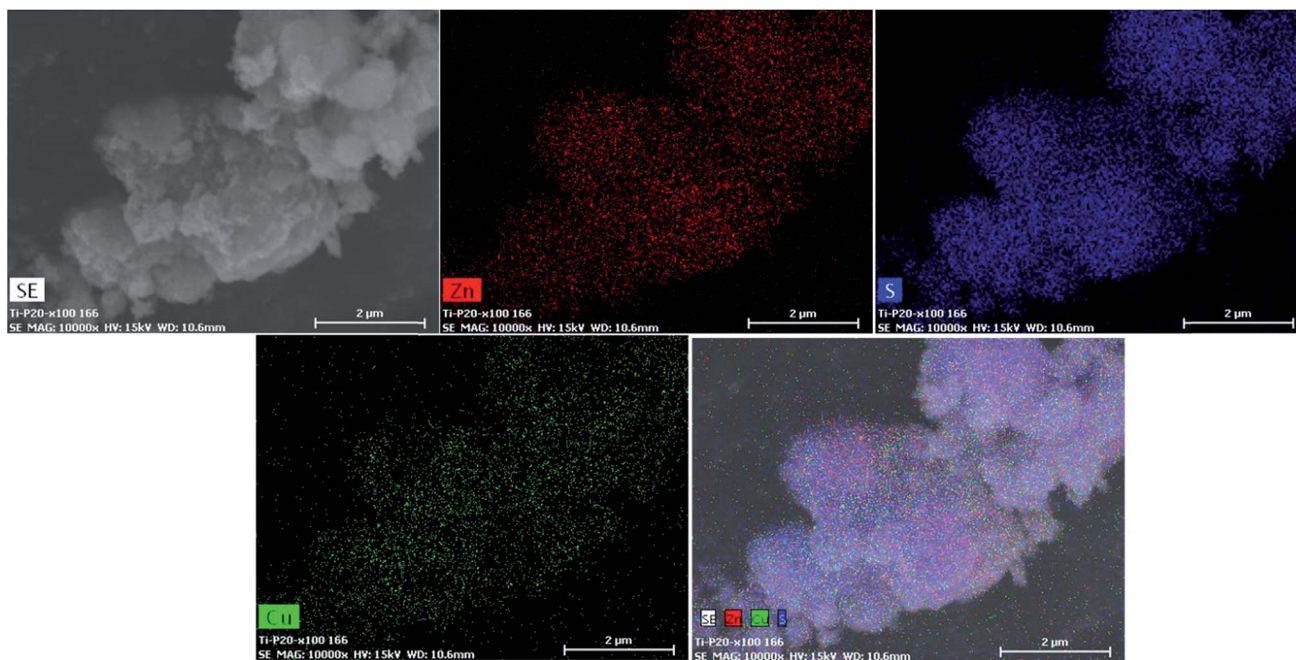


Fig. 5 EDS elemental mapping of the ZC2 sample.



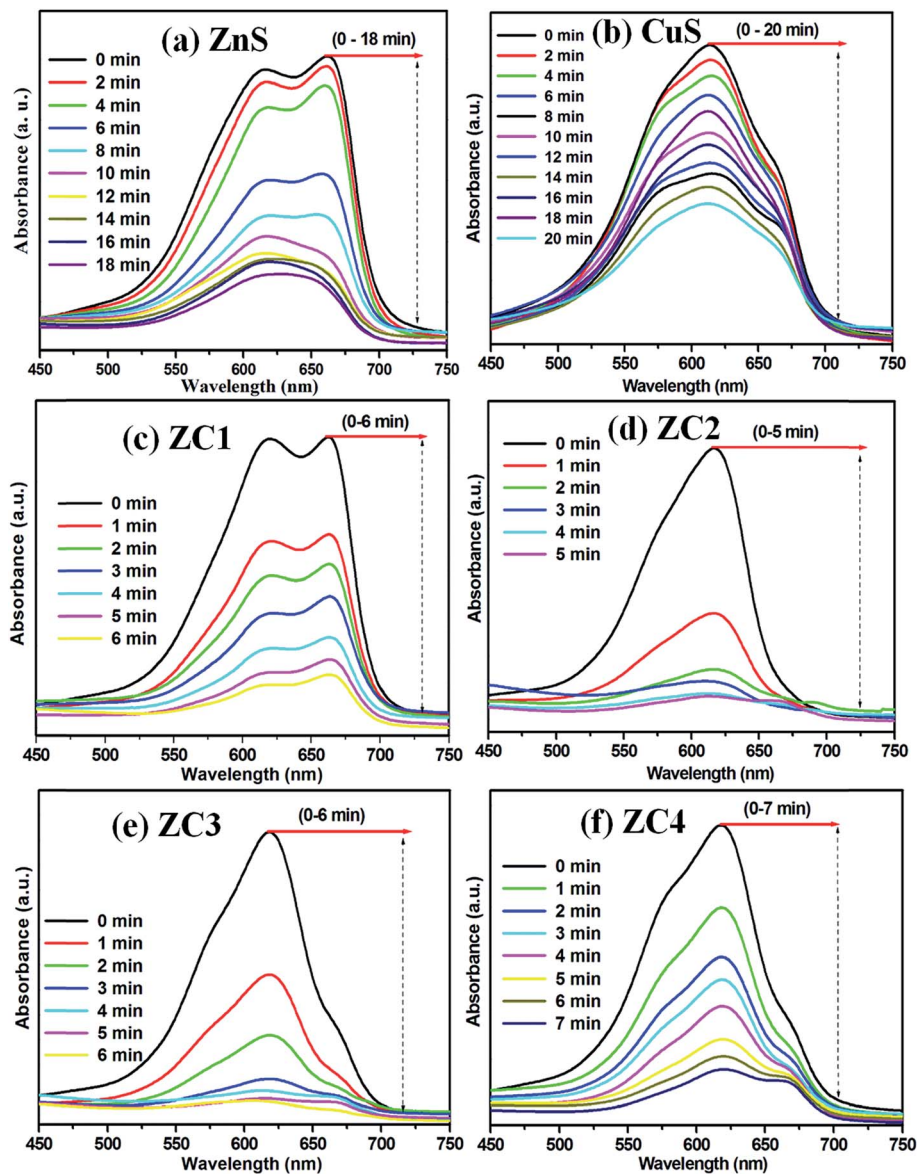


Fig. 6 UV absorption spectra of MB (a) ZnS, (b) CuS, (c) ZC1, (d) ZC2, (e) ZC3 and (f) ZC4.

length. When the CuS concentration was 0.025 M, the ZnS nanoparticles were partially coated on the surface of the CuS nanorods, as shown in Fig. 4(c1)–(c3). As the CuS concentration was increased to 0.050 (ZC2), the surface of the nanorods was fully covered by ZnS nanoparticles, as shown in Fig. 4(d1)–(d3). With further increases in the concentration of CuS to 0.075 (ZC3) and 0.1 M (ZC4), the CuS nanorods were embedded in the ZnS nanostructures, as shown in Fig. 4(e1)–(f3). The high-resolution TEM (HRTEM) images obtained from the edge of individual nanorods revealed the formation of heterojunctions of ZnS and CuS nanoparticles. Fig. 5 shows an element mapping analysis of Zn, Cu and S for the ZC2 sample. The signals for Zn, Cu and S were uniform, indicating the homogeneous distribution of Zn, S and Cu elements in the CuS/ZnS composites.

MB photodegradation is of great significance for environmental pollutant treatment, and it is commonly used as a model

pollutant for evaluating the degradation efficiency of a photocatalyst. The time-dependent UV-Vis absorption spectrum of CuS/ZnS composites under visible light irradiation is shown in Fig. 6(a)–(f). Fig. 6(a) and (b) show the time-dependent UV absorption spectra of bare ZnS and CuS, respectively. MB was completely decomposed in 20 and 18 min, respectively. Fig. 6(c)–(f) show the photodegradation of samples ZC1–ZC4, respectively. The initial absorbance of the peak disappeared completely after 5 min of irradiation for sample ZC2, as shown in Fig. 6(d). As the lower concentration of the CuS is 0.025 M, the photodegradation time increased to 6 min, as shown in Fig. 6(c), and when the concentration increased to 0.075 and 0.1 M, the degradation time increased to 6 and 7 min, as shown in Fig. 6(e) and (f), respectively. It can be clearly seen that the photocatalytic activities of the CuS/ZnS composites depended on the amount of CuS concentration. The highest photocatalytic



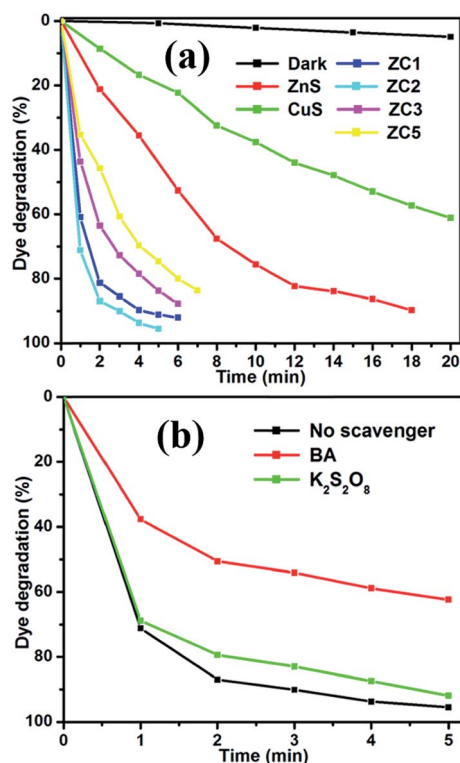


Fig. 7 Effect of dye degradation efficiency (a) time (min) vs. dye degradation (%), and (b) MB over CuS/ZnS in the presence of various scavengers under visible light irradiation.

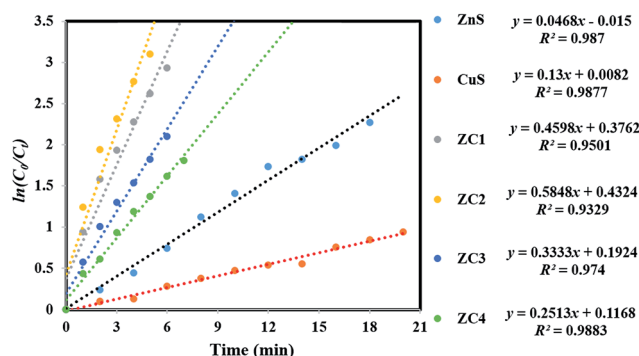


Fig. 8 Plots of $\ln(C_0/C_t)$ as a function of time (min) for the photo-degradation of MB over the CuS/ZnS nanocomposites.

activity was obtained for the 0.025 M sample (ZC2). Fig. 7(a) shows the blank experiment, carried out in the absence of a catalyst, with no change in the degradation of MB. This observation clearly shows that the degradation of MB occurred only with the photocatalyst present. The degradation percentages of MB in the presence of ZnS, CuS, ZC1, ZC2, ZC3 and ZC4 were 89.70, 61.05, 92.07, 95.51, 87.79 and 83.65%, respectively. The photocatalytic activity strongly depends on the initial concentration of Cu²⁺ in the ion-exchange reaction. The photocatalytic reaction of the pure sample is decreased and it is drastically increased after the incorporation of CuS in the composition. The addition of CuS formed a heterojunction,

Table 1 Observed pseudo-first order rate constants, R^2 values, maximum degradation (%) and time required for maximum degradation of CuS/ZnS nanostructures

Sample	K_{app} (CuS/ZnS)	R^2	Maximum degradation (%)	Time taken for maximum degradation (min)
ZnS	0.0046	0.9870	89.70	20
CuS	0.1300	0.9877	61.05	18
ZC1	0.4598	0.9501	92.07	06
ZC2	0.5848	0.9329	95.51	05
ZC3	0.3333	0.9740	87.79	06
ZC4	0.2513	0.9883	83.65	07

resulting in a significant increase in the photocatalytic activity. Sample ZC2 shows the highest photocatalytic activity, and a further increase in the concentration of CuS leads to a decrease in the activity. The increased concentration of CuS may lead to a decrease in the interface and act as recombination centres of electron-hole pairs. Thus, the separation efficiency of the electron-hole pairs decreases with a decrease in the activity. Many results reported the same phenomena, such as ZnS/CuS nanotubes, CuS/ZnS nanowires, Cu-ZnS, TiO₂/ZnS and ZnO/SnO₂.^{50–54}

Different reactive species, such as the hydroxyl radical ($\cdot\text{OH}$) and superoxide radical ion ($\text{O}_2^{\cdot-}$), are generated in the photocatalysis process. The main reactive species are electrons and holes. To understand the photocatalysis process in-depth, different scavenging environments were studied. From the literature, it is apparent that benzoic acid and potassium persulfate can act as good electron and hole scavengers.^{55,56} Benzoic acid (BA) is used to trap electrons of the conduction band and potassium persulfate ($\text{K}_2\text{S}_2\text{O}_8$) is used to trap holes of the valence band before the inclusion of the CuS/ZnS photocatalyst. Fig. 7(b) shows the degradation of methylene blue in the presence of the hole scavenger, $\text{K}_2\text{S}_2\text{O}_8$. In the presence of the $\text{O}_2^{\cdot-}$ scavenger, only 5% was decreased compared with the scavenger free photocatalyst for MB. To examine the reactive species, another experiment was performed using BA as the $\cdot\text{OH}$ scavenger. As a result, only a 62% MB degradation was observed instead of 95% for the scavenger free photocatalyst. This indicates that the superoxide radical was the most dominant species. Due to the scavenging of the superoxide radical, the probability of electron-hole pair recombination diminishes and the rate of reaction increases.^{57,58}

The kinetics of the MB photodegradation of all the samples is shown in Fig. 8. The catalytic rate constants of the MB degradation by bare ZnS, CuS and CuS/ZnS nanocomposites are shown in Table 1, which are calculated by the linear fitting of $\ln(C_0/C_t)$ versus irradiation time.⁵⁹ The apparent rate constant (K_{app}) value was determined for the photocatalyst. In the case of bare ZnS and CuS, the value of K_{app} was 0.0046 and 0.1300 min^{-1} , respectively. Whereas, when the composite of ZnS/CuS was used as the photocatalyst, the value of K_{app} increased to 0.4598, 0.5848, 0.3333 and 0.2513 min^{-1} for ZC1, ZC2, ZC3 and ZC4, respectively. The K_{app} value increased from 0.4598 to 0.5848 min^{-1} with the addition of CuS concentrations from



Table 2 Comparison of photocatalytic performance of recently reported ZnS composites and this work

S. no.	Material	Morphology	Illumination source	Target dye	Dye deg. (%)	Time (min)	Ref. no.
1	ZnS-GO	Nanosheets	Sunlight	Methylene blue	86	150	67
2	GQD/ZnS	Nanobelts	Xenon lamp	Rhodamine B	—	40	68
3	ZnS/CdS/Ag ₂ S	Nanoparticles	Sunlight	Congo dye	98	120	69
4	CdS/ZnS	Quantum dots	Xenon lamp	Methylene blue	88	180	70
5	ZnO/ZnS _x Se _{1-x}	Nanowire	Xenon lamp	Methylene blue	—	180	71
6	ZnO/ZnS	Hollow structures	UV lamp	<i>p</i> -Chlorophenol	90	120	72
7	ZnS/GAs	3D nanosheets	Xenon lamp	Methylene blue	98	60	73
8	ZnO/ZnS	Nanowire	Hg-arc lamp	Methylene orange	92	40	74
9	Zn _x Cd _{1-x} S/rGO	Layered structure	Xenon lamp	Methylene blue	98	120	75
10	CuS/ZnS	Nanorods	Xenon lamp, MAX-303, Asahi spectra	Methylene blue	95	05	This work

0.025 M to 0.050 M. When the CuS concentration was increased from 0.075 M to 0.1 M, the K_{app} value decreased from 0.5848 to 0.2513 min⁻¹. The kinetic data obtained by the pseudo-first-order model, such as K_{app} , the corresponding correlation coefficients (R^2) and the maximum dye degradation in the presence of CuS/ZnS composites are presented in Table 1. Comparison of photocatalytic performance was made between the materials developed in this work and other reported ZnS composites, as shown in Table 2. The CuS/ZnS composites showed 95.51% degradation in 5 min, which was higher than any other reported materials, indicating the enhanced photocatalytic activity under visible light illumination.

Based on the above results, the photocatalytic mechanism for the CuS/ZnS heterojunction can be explained as follows.^{60,61} Upon visible light irradiation, CuS is excited and the electron-hole pairs are produced. According to the interfacial charge transfer (IFCT) mechanism,⁶² the electrons in CuS migrate to the conduction band (CB) of ZnS.^{63,64} This transfer process is thermodynamically favourable because the CB of CuS is higher than that of ZnS after equilibrium, as shown in Fig. 9. Meanwhile, the photogenerated holes that remain in the valence band of CuS are positioned at much higher values than those of the VB of ZnS. Therefore, the holes produced within the material cannot migrate to ZnS, and therefore did not oxidize H₂O.^{65,66}

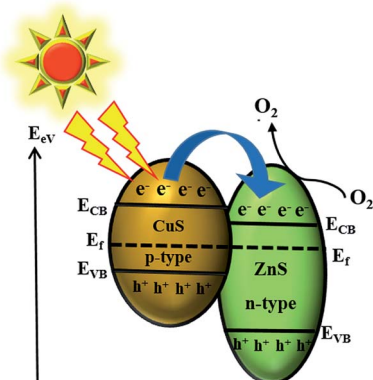


Fig. 9 Schematic mechanism of electron-hole pair separation in the p-type CuS/n-type ZnS heterojunction.

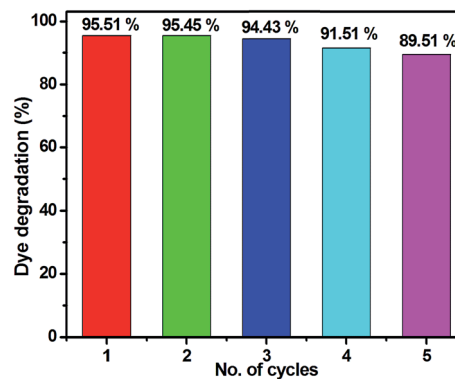


Fig. 10 Reusability of the photocatalyst.

The reusability of the ZC2 photocatalyst was analysed to detect the stability of the material. After the photocatalysis experiment, the catalyst was separated from the reaction mixture by centrifugation. As shown in Fig. 10, the catalyst did not exhibit any significant loss after five cycles of the degradation of MB under the same conditions. The obtained degradation values were 95.51, 95.45, 94.43, 91.51 and 89.51% for the first, second, third, fourth and fifth cycles, respectively. The results confirm that the obtained CuS/ZnS photocatalyst shows good stability and sustainability. In addition to that, we have investigated the structural property of the used catalyst (ZC2) by XRD and XPS analysis, as shown in Fig. S1 (ESI†). It can be clearly observed that the phase and structure of the CuS/ZnS composite remained unchanged after the photocatalytic cycles, suggesting that the sample is stable in the present photocatalytic degradation process. The photocatalytic structural stabilities were also investigated by XPS spectra. From the spectra, it is clear that the binding energies of Zn 2p, Cu 2p and S 2s of CuS/ZnS showed no peak-shift compared to those of the fresh sample.

4. Conclusion

ZnS nanoparticles were decorated on the surface of CuS nanorods by a hydrothermal method. This methodology led to the formation of heterojunctions between p-type



ZnS. The investigation of photocatalytic activity indicated that the CuS/ZnS possessed a higher photocatalytic efficiency which was 3 times faster than that of pure ZnS and CuS. The photocatalytic activity depended on the content of CuS in the CuS/ZnS composites. The 0.050 M CuS showed the highest photocatalytic activity for the MB degradation. Photogenerated superoxide (O_2^-) radicals over CuS/ZnS were the main reactive species for the degradation of MB. The obtained material is useful for environmental remediation in degrading organic pollutants.

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